

## General Chemistry 1

CREDIT	3	INSTRUCTOR	S. Carroll Brooks III
OFFICE	N307 Baekyanggwan	OFFICE HOURS	By appointment
TIME	13:20 ~ 15:00	CLASSROOM LOCATION	TBA
E-MAIL	<a href="mailto:brookssc@yonsei.ac.kr">brookssc@yonsei.ac.kr</a>		

### [COURSE INFORMATION]

COURSE DESCRIPTION & GOALS	This course focuses on the science of matter, its physical properties and composition, as well as how it changes. The topic are organized into three parts: basic concepts, atoms and molecules, and states of matter.
PREREQUISITE	High School Chemistry recommended
COURSE REQUIREMENTS	Students should arrive a few minutes early to class each day with their cell phones turned off and a pen or pencil and paper or printed out copy of the PowerPoint course material. Successful students will incorporate readings from the text, participation in class, in-chapter problems, end-of-chapter problems, the textbook website resources, and independent study outside of class.
GRADING POLICY	Attendance 10%, Participation 10%, Midterm Assessment 40%, Final Assessment 40%.
TEXTS & NOTES	Kotz, Treichel, Townsend, and Treichel: Chemistry & Chemical Reactivity 9e, Cengage Learning
INSTRUCTOR'S PROFILE	Ph.D. Cornell University B.A. Wayne State University Associate Professor Department of Life Science and Biotechnology and Underwood International College

[WEEKLY SCHEDULE]

WEEK (PERIOD)	WEEKLY TOPIC & CONTENTS	COURSE MATERIAL & ASSIGNMENTS	REFERENCE
1	<p>PART I: CONCEPTS OF CHEMISTRY.</p> <p>1. Basic Concepts of Chemistry.</p> <p>Let's Review: The Tools of Quantitative Chemistry.</p> <p>2. Atoms, Molecules, and Ions.</p>		
2	<p>3. Chemical Reactions.</p> <p>4. Stoichiometry: Quantitative Information from Chemical Reactions.</p>		
3	<p>5. Principles of Chemical Reactivity: Energy and Chemical Reactions.</p>	Midterm Assessment	
4	<p>PART II: ATOMS AND MOLECULES.</p> <p>6. The Structure of Atoms.</p> <p>7. The Structure of Atoms and Periodic Trends.</p>		
5	<p>8. Covalent Bonding and Molecular Structure.</p> <p>Part III: STATES OF MATTER.</p> <p>10. Gases and Their Properties.</p>		
6	<p>11. Intermolecular Forces and Liquids.</p> <p>13. Solutions and Their Behavior.</p>	Final Assessment	